

VEATCHITE, A NEW CALCIUM BORATE FROM LANG, CALIFORNIA

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ABSTRACT

Veatchite, $\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 2\text{H}_2\text{O}$, is a new mineral from Lang, California. It is monoclinic, with lattice constants $a_0 = 6.72\text{\AA}$, $b_0 = 41.26\text{\AA}$, $c_0 = 41.20\text{\AA}$, $\beta = 67^\circ$. Perfect $\{010\}$ and imperfect $\{001\}$ cleavages; $H = 2$, $G = 2.69$. Biaxial, positive; $Z:c[001] = -38^\circ$, $Y = b[010]$; indices (Na): $n_X = 1.551$, $n_Y = 1.553$, $n_Z = 1.621$, all ± 0.002 ; $2V = 37^\circ$. Occurs in white cross-fibre veins in limestone and howlite.

INTRODUCTION

The mineral here described was found by Mr. William H. Nisson, of Petaluma, California, at the old colemanite mine at Lang, Los Angeles County, California, in September 1936. Veatchite, as the new species will be named, occurs in white cross-fibre veins up to 6 mm. across, with dark grey clay, cutting massive white limestone and howlite. The new species is apparently rare, since only a few small specimens have been found.

MORPHOLOGY

No terminated crystals of veatchite were found. The vein fibres are rather coarse (up to 1 mm. across) and exhibit two cleavages. The better cleavage is $b\{010\}$. The other cleavage is at an angle of $67^\circ (\pm 1^\circ)$ to the length of the fibres and is taken as $c\{001\}$. The symmetry is thus probably monoclinic. The perfect $b\{010\}$ cleavage produces very thin plates tabular parallel to $b\{010\}$ and elongated parallel to the length of the fibres, terminated at either end by the imperfect basal cleavage. The fibres form subparallel aggregates, and it is only by cleaving down to a very thin plate that a single individual can be obtained.

PHYSICAL PROPERTIES

The cleavage of veatchite is perfect and easy parallel to $b\{010\}$, and imperfect parallel to $c\{001\}$. The hardness is 2. The specific gravity determined by suspension in bromoform is 2.69. It is clear and colorless with a silky luster.

OPTICAL PROPERTIES

	$n(\text{Na})$	
X	1.551	} ± 0.002 Positive $2V = 37^\circ (\pm 2^\circ)$ Dispersion $r > v$, perc.
$Y = b[010]$	1.553	
$Z:c[001] = -38^\circ$	1.621	

X-RAY MEASUREMENTS

Rotation photographs were taken on cleavage fragments with both $c[001]$ and $a[100]$ as axes of rotation. A Weissenberg photograph was taken with $c[001]$ as the rotation axis, the zero and first layer lines appearing on the same picture because of the extremely long translation distance along the c axis. These photographs gave a monoclinic unit cell with the dimensions:

$$a_0 = 6.72 \text{ \AA}, b_0 = 41.26 \text{ \AA}, c_0 = 41.20 \text{ \AA}, \beta = 67^\circ \text{ (from morphology)}$$

$$\text{Axial ratio: } a_0 : b_0 : c_0 = 0.163 : 1 : 0.998.$$

Using the unit cell dimensions given above and a density of 2.69, the molecular weight of the unit cell becomes 17,126.

The space group was not determinable because of the uncertainty of indexing all of the spots due to the large dimensions of the unit cell in the directions c and b .

COMPOSITION

Table 1 gives the chemical analysis and the atomic content of the unit cell using the molecular weight 17,126.

TABLE 1. ANALYSIS OF VEATCHITE AND CONTENT OF UNIT CELL

	1.	2.	3.	4.	5.	6.
CaO	31.08	31.20	0.5563 = 2 × 0.2782	Ca 95.4	96	CaO 31.41
B ₂ O ₃	59.03	59.26	0.8510 = 3 × 0.2873	B 291.4	288	B ₂ O ₃ 58.51
H ₂ O	9.50	9.54	0.5300 = 2 × 0.2651	O 623.3	624	H ₂ O 10.08
Insol.	0.62			H 181.6	192	
	100.23	100.00				100.00

1. Analysis by F. A. Gonyer.
2. Analysis calculated to 100%.
3. Molecular ratios.
4. Atomic content of the unit cell.
5. Theoretical unit cell content for 48(Ca₂B₆O₁₁ · 2H₂O).
6. Calculated composition for Ca₂B₆O₁₁ · 2H₂O.

The calculated density using the theoretical unit cell content is 2.665.

Pyrognostics. Before the blowpipe, on charcoal, veatchite fuses at 1.5 with intumescence to an opaque white bead. It gives the characteristic greenish yellow boron flame. In the closed tube it yields a moderate amount of water, which gives a strong acid reaction.

THE SYSTEM R₂B₆O₁₁ · nH₂O'

The formula Ca₂B₆O₁₁ · 2H₂O for veatchite places it in the same chemical system with five other calcium and magnesium borates. Table 2 tabulates some of the properties of this system in order of increasing degree of hydration.

TABLE 2

Name	Composition	Density	Mean Ref. Index	Crystal System	Cleavage
Veatchite	$\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 2\text{H}_2\text{O}$	2.69	1.586	Mono.	{010} per. {001} imp.
Colemanite	$\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 5\text{H}_2\text{O}$	2.42	1.600	Mono.	{010} per. {001} dis.
Hydroboracite	$\text{CaMgB}_6\text{O}_{11} \cdot 6\text{H}_2\text{O}$	2.17	1.548	Mono.	{010} per. {001} per.
Meyerhofferite	$\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 7\text{H}_2\text{O}$	2.12	1.530	Tri.	{010} per.
Inyoite	$\text{Ca}_2\text{B}_6\text{O}_{11} \cdot 13\text{H}_2\text{O}$	1.87	1.510	Mono.	{001} per.
Inderite ¹	$\text{Mg}_2\text{B}_6\text{O}_{11} \cdot 15\text{H}_2\text{O}$	1.80	1.490		

It is noteworthy in the above table that the specific gravity becomes less with increasing degree of hydration; that there is a more or less systematic change in the mean index of refraction; and that all the members of the system have at least one perfect cleavage, usually $b\{010\}$.

OCCURRENCE

The minerals of the system $\text{R}_2\text{B}_6\text{O}_{11} \cdot n\text{H}_2\text{O}$ occur intimately associated in deposits of a similar nature in several California localities, along with other borates, usually ulexite ($\text{NaCaB}_5\text{O}_9 \cdot 8\text{H}_2\text{O}$) and howlite ($\text{H}_5\text{Ca}_2\text{B}_5\text{SiO}_{14}$). Colemanite is always the most abundant mineral. Such deposits are found at Lang, Borate, several localities in Ventura County, and in the Ryan and Mt. Blanco districts near Death Valley. They are all interbedded with lake bed deposits of Tertiary age.²

The exact mode of occurrence of veatchite at Lang is not known. Although the specimens show veatchite veins cutting limestone and howlite, they were found as part of a small boulder lying in the wash, and none of the material has yet been found in place.

Name. It seems fitting to name this new borate *veatchite* after Dr. John A. Veatch, who was the first to detect the presence of borates in the mineral waters of California, on January 8, 1856, as related in a letter from Dr. Veatch to Henry G. Hanks, in the California State Mining Bureau, Third Annual Report of the State Mineralogist (1883).

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¹ Godlevsky, M. N., Mineralogical Investigation of the Inder borate deposit: *Mem. Soc. Russ. Mineral.*, 2d Ser., vol. 66, No. 2, pp. 315-344, 1937.

² Foshag, W. F. The origin of colemanite deposits: *Econ. Geol.*, vol. 16, p. 203, 1921.