very useful while doing geologic mapping in remote areas. During mapping, the staining has been carried out on individual specimens during traverses or, more commonly, several specimens were stained at the end of a day at a base camp.

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### THE CRYSTAL STRUCTURE OF FLUELLITE: A CORRECTION

BRIAN B. GUY, G. A. JEFFREY, Crystallography Laboratory, University of Pittsburgh, Pittsburgh, Pennsylvania and R. Van Tassel, Institut Royale des Sciences Naturelles de Belgique, Bruxelles, Belgium.

In a recent X-ray single-crystal study of a specimen of the mineral fluellite, Al<sub>2</sub>PO<sub>4</sub>F<sub>2</sub>OH·7H<sub>2</sub>O, from Cornwall, England, Guy and Jeffrey (1966) determined the presence of the phosphate ion in the structure, although some previous investigations had suggested the composition was that of an aluminum hydroxy-fluoride monohydrate. Two publications, one by Van Tassel (1959) and the other published by the USSR Academy of Science (1963), also indicating the presence of phosphate in this mineral, have been noted by one of us (RVT).

Since reference to these publications was overlooked in the crystal structure paper, a summary of the chemical composition data from them is presented in Table 1.

## REFERENCES

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TABLE 1

Author	Specimen locality	Method and results of analysis	Proposed formulae
Van Tassel (1959)	Argenteau, Belgium	Chemical Analysis (recalculated) Al: 18.8 F: 16.0 PO4: 27.6	3Al <sub>2</sub> O <sub>3</sub> ·4AlF <sub>3</sub> ·2P <sub>2</sub> O <sub>5</sub> ·27H <sub>2</sub> O (from Van Tassel, no OH content)
		OH: 6.5 OH calculated value to balance Al charge.  H <sub>2</sub> O: 31.1 H <sub>2</sub> O by difference. $D_{\rm m} = 2.12$ . Powder data for material from Cornwall, Bavaria, and Belgium.	Al <sub>2,4</sub> PO <sub>4</sub> F <sub>2,9</sub> (OH) <sub>1,8</sub> ·6H <sub>2</sub> O (calculated, assuming Al balanced by OH)
Chukrov (1963)	Kazakhastan	Chemical Analysis: Al: 16.83 17.04 F: 14.80 14.62 PO <sub>4</sub> : 28.44 28.97 OH: 3.32 3.59 OH calculated to $H_2O: 35.89 35.36$ balance Al charge, etc: 0.77 0.33 $D_m = 2.17$ (both specimens) Powder data for Kazakhastan material.	Al <sub>2</sub> (F·OH) <sub>3</sub> PO <sub>4</sub> ·7H <sub>2</sub> O (may be 6 or 7 H <sub>2</sub> O per formula unit. F:OH is 4:1)

GUY, B. B. AND G. A. JEFFREY (1966) The Crystal Structure of Fluellite, Al<sub>2</sub>PO<sub>4</sub>F<sub>2</sub>(OH) ·7H<sub>2</sub>O. Amer. Mineral. 51, 1579-1592.

Van Tassel, R. (1959) Autunite, apatite, delvauxite, évansite et fluellite de la région de Visé. Bull. Soc. belge Géol., 68, 226–248.

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# BARIUM-VANADIUM MUSCOVITE AND VANADIUM TOURMALINE FROM MARIPOSA COUNTY, CALIFORNIA: A CORRECTION

Kenneth G. Snetsinger, NASA Ames Research Center, Moffet Field, California.

Dr. L. G. Berry has drawn my attention to the absence of the (006) and (024) reflections from the X-ray data for barium-vanadium muscovite in my paper (Amer. Mineral. 51, 1623–1639); both reflections are typical of  $2M_1$  muscovite. I neglected to include (006) in the data; it has d (meas.) = 3.339 Å, and an intensity of 90. These values were obtained from film because in diffractometer work the  $2\theta$  range in question was swamped by the intense internal standard quartz peak; d (calc.) of (006) is 3.336 Å. I find no indication of (024) on the X-ray film of the mica; (024) is considerably less intense than nearby (006) (cf. ASTM card 6-0263), and presumably is masked by (006).

#### ERRATA

Nester, J. F. (1967) Growth of synthetic calcite crystals. **52,** 276–280: p. 276, for "Li<sub>2</sub>O<sub>3</sub>" read "La<sub>2</sub>O<sub>3</sub>."