

TABLE 1

Author	Specimen locality	Method and results of analysis	Proposed formulae
Van Tassel (1959)	Argenteau, Belgium	Chemical Analysis (recalculated) Al: 18.8 F: 16.0 PO ₄ : 27.6 OH: 6.5 } OH calculated value to } balance Al charge. H ₂ O: 31.1 } H ₂ O by difference. <i>D_m</i> = 2.12. Powder data for material from Cornwall, Bavaria, and Belgium.	3Al ₂ O ₃ · 4AlF ₃ · 2P ₂ O ₅ · 27H ₂ O (from Van Tassel, no OH content) Al _{2.4} PO ₄ F _{2.9} (OH) _{1.3} · 6H ₂ O (calculated, assuming Al balanced by OH)
Chukrov (1963)	Kazakhstan	Chemical Analysis: Al: 16.83 17.04 F: 14.80 14.62 PO ₄ : 28.44 28.97 OH: 3.32 3.59 } OH calculated to H ₂ O: 35.89 35.36 } balance Al charge. etc: 0.77 0.33 <i>D_m</i> = 2.17 (both specimens) Powder data for Kazakhstan material.	Al ₂ (F · OH) ₃ PO ₄ · 7H ₂ O (may be 6 or 7 H ₂ O per formula unit. F:OH is 4:1)

GUY, B. B. AND G. A. JEFFREY (1966) The Crystal Structure of Fluellite, Al₂PO₄F₂(OH) · 7H₂O. *Amer. Mineral.* **51**, 1579-1592.

VAN TASSEL, R. (1959) Autunite, apatite, delvauxite, évanosite et fluellite de la région de Visé. *Bull. Soc. belge Géol.*, **68**, 226-248.

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BARIUM-VANADIUM MUSCOVITE AND VANADIUM TOURMALINE FROM MARIPOSA COUNTY, CALIFORNIA: A CORRECTION

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Dr. L. G. Berry has drawn my attention to the absence of the (006) and (024) reflections from the X-ray data for barium-vanadium muscovite in my paper (*Amer. Mineral.* **51**, 1623-1639); both reflections are typical of 2M₁ muscovite. I neglected to include (006) in the data; it has *d* (meas.) = 3.339 Å, and an intensity of 90. These values were obtained from film because in diffractometer work the 2θ range in question was swamped by the intense internal standard quartz peak; *d* (calc.) of (006) is 3.336 Å. I find no indication of (024) on the X-ray film of the mica; (024) is considerably less intense than nearby (006) (*cf.* ASTM card 6-0263), and presumably is masked by (006).

ERRATA

NESTER, J. F. (1967) Growth of synthetic calcite crystals. **52**, 276-280: p. 276, for "Li₂O₃" read "La₂O₃."