

Transformation of 2-line ferrihydrite to 6-line ferrihydrite under oxic and anoxic conditions

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ABSTRACT

Mineralogical transformations of 2-line ferrihydrite were studied under oxic and Fe³⁺-reducing conditions to establish the role, if any, of 6-line ferrihydrite (“well” organized ferrihydrite) in the reaction pathway and as a final product. In oxic experiments, concentrated suspensions (0.42 mol/L Fe³⁺ in 0.1 mol/L NaClO₄) of freshly synthesized 2-line ferrihydrite, with and without 3% Ni²⁺, were aged at an initial pH = 7.2 (unbuffered and unadjusted) and 25 °C for more than three years. X-ray diffraction, transmission electron microscopy, and Mössbauer spectroscopy measurements were performed on the solids after different aging periods. The primary mineralogical products observed were 6-line ferrihydrite and goethite, with minor hematite. Aggregation and crystallization of the 2-line ferrihydrite liberated protons and depressed suspension pH, but coprecipitated Ni²⁺ retarded this process. The joint, interrelated effects of Ni and pH influenced both the extent of conversion of 2-line ferrihydrite and the identity of the major transformation products. Six-line ferrihydrite dominated in the Ni ferrihydrite suspension, whereas goethite dominated in the absence of Ni. Aggregation-induced crystallization of 2-line ferrihydrite particles seemed responsible for 6-line ferrihydrite formation. Mineralogical changes to Ni ferrihydrite under anaerobic conditions were investigated at circumneutral pH using the Fe³⁺-reducing bacterium *Shewanella putrefaciens*. Residual 6-line ferrihydrite dominated bioreduced samples that also contained goethite and magnetite. The conversion of 2-line ferrihydrite to 6-line ferrihydrite was considerably more rapid under anaerobic conditions. The sorption of biogenic Fe²⁺ apparently induced intra-aggregate transformation of 2-line ferrihydrite to 6-line ferrihydrite. Collectively, abiotic and biotic studies indicated that 6-line ferrihydrite can be a transformation product of 2-line ferrihydrite, especially when 2-line ferrihydrite is undergoing transformation to more stable hematite or magnetite.

INTRODUCTION

Poorly crystalline Fe³⁺-oxides (ferrihydrites) that commonly exist in soils and sediments are thermodynamically unstable and, with time, transform to more crystalline Fe³⁺ oxides, e.g., goethite, hematite (Cornell and Schwertmann 1996). The transformation into crystalline Fe³⁺ oxides is widely believed to proceed via two competing mechanisms (Feitknecht and Michaelis 1962; Schwertmann and Murad 1983). Goethite formation involves the dissolution of ferrihydrite followed by nucleation and precipitation of the crystalline oxide in solution, whereas hematite forms through dehydration and internal atomic arrangement of the solid ferrihydrite, with less dependence on dissolution mechanism. The predominant mechanism depends primarily on suspension pH (Schwertmann and Murad 1983), with hematite formation achieving a maximum near pH 8 and goethite near pH 4 and 11.

Aggregation-based crystal growth and phase transformation has also been advanced as an alternative mechanism to explain the transformation of 2- to 3-nm ferrihydrite particles to polycrystalline Fe³⁺ oxides at pH conditions where the Fe-oxide phase is relatively insoluble (Banfield et al. 2000). Moreover,

Schwertmann et al. (1999) suggested that nano-particle aggregation was essential to induce crystallization of ferrihydrite to hematite. Ferrihydrite aggregation is maximal near the pH of zero net charge of the ferrihydrite “surface” (pH ~8), and crystallizing hematite nuclei are apparently supplied by a short-range dissolution process involving precursor aggregates.

Synthetic and natural ferrihydrites are poorly ordered, but both exhibit a continuum in structure from amorphous to partly crystalline (Carlson and Schwertmann 1981; Cornell and Schwertmann 1996). Ferrihydrite exhibits a range of XRD patterns; the least crystalline variety exhibits two broad peaks (2-line ferrihydrite),¹ and the more crystalline variety exhibits six broad peaks (6-line ferrihydrite). Several structural models of 6-line ferrihydrite have been proposed, with the defective hematite structural model proposed by Chukrov et al. (1973) and Towe and Bradley (1967) being the most widely accepted. In contrast, Drits et al. (1993) proposed on the basis of XRD simulations that all natural and synthetic ferrihydrites are multi-component phases comprised of defect-free and defective ferrihydrite mixed with ultradisperse hematite. The main dif-

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¹The terms two-line ferrihydrite and HFO (hydrrous ferric oxide) are commonly used synonymously (Cornell and Schwertmann 1996). The term HFO (e.g., Dzombak and Morel 1990) is applied to a material synthesized in the laboratory by rapid hydrolysis of a Fe³⁺ salt solution, with approximately 4–8 h aging at pH 7.

ference between 2-line ferrihydrite and 6-line ferrihydrite is the size of their coherent scattering domains.

Cornell and Schwertmann (1996) and Schwertmann et al. (1999) suggested that dilute suspensions of 2-line ferrihydrite do not transform to 6-line ferrihydrite with time, because these two forms of ferrihydrite precipitate under different conditions (e.g., Schwertmann and Cornell 1991; Schwertmann et al. 1999). In contrast, we recently observed the partial transformation of an aged 2-line Ni ferrihydrite (Ni/[Ni + Fe] mole fraction of ~0.03) to 6-line ferrihydrite in the presence of a dissimilatory iron-reducing bacterium (DIRB; *Shewanella putrefaciens*, strain CN32) at 25 °C in bicarbonate-buffered solution at circumneutral pH (Fredrickson et al. 2001). Our finding of 6-line ferrihydrite as a biotransformation product was consistent with the recently observed conversion of 2-line arsenate (As) ferrihydrite to 6-line As ferrihydrite in sediment (Rancourt et al. 2001).

Several authors have shown that 6-line ferrihydrite may convert to hematite with heat treatment (Johnston and Lewis 1983; Stanjek and Weidler 1992; Weidler 1995). Implied but not documented is that 6-line ferrihydrite may be an intermediate or metastable structure in the conversion of 2-line ferrihydrite to hematite. Recently, Schwertmann et al. (1999) attributed the transformation of an aqueous suspension of 6-line ferrihydrite to hematite, without any transitional phases, to crystallization within ferrihydrite aggregates.

In the present work, experiments were performed to evaluate whether 2-line ferrihydrite transforms to 6-line ferrihydrite during conversion to more crystalline Fe oxides such as hematite and magnetite. Our objectives were to define: (1) whether 2-line ferrihydrite gradually crystallizes to a more organized phase (e.g., 6-line ferrihydrite) in aerobic and anaerobic environments, (2) the effect of small amounts of coprecipitated Ni²⁺ in ferrihydrite on such transformation, and (3) the commonality of 6-line ferrihydrite as an intermediate transformation product. These objectives were resolved by first performing aging experiments with concentrated suspensions (0.42 mol/L) of freshly precipitated 2-line and 2-line Ni ferrihydrites in 0.1 mol/L NaClO₄. These were aged at 25 °C and pH 7.2 (initial) for 41 and 38 months, respectively. The aging conditions were chosen to promote the initial aggregation of ferrihydrite, and hence, hematite formation (Schwertmann and Fischer 1966; Fischer and Schwertmann 1975; Cornell and Giovanoli 1985; Schwertmann et al. 1999). The resulting mineral transformation products were characterized using a variety of techniques. Second, low-temperature Mössbauer measurements were performed on biotransformed Ni ferrihydrite from two incubations of Fredrickson et al. (2001) that produced different mineral associations with residual ferrihydrite. The combined results of these two experimental activities demonstrated that 6-line ferrihydrite can form as a transformation product of 2-line ferrihydrite under various conditions.

MATERIALS AND METHODS

Materials

The 2-line ferrihydrite and 2-line Ni ferrihydrite coprecipitates were prepared as described by Ainsworth et al. (1994) using Fe(NO₃)₃·9H₂O and

Ni(NO₃)₂·6H₂O. The suspension pH was brought to 7.2 over the course of approximately 2 h using anoxic, CO₂-free 2 mol/L NaOH. The suspension pH remained almost constant during the 24 h after precipitation. For the Ni ferrihydrite, the mole proportions of Fe and Ni in solution before precipitation were 0.95 and 0.05, respectively. The Fe ferrihydrite and Ni ferrihydrite were washed with anoxic, CO₂-free 0.1 mol/L NaClO₄ to remove nitrate, whose nitrate abundances were <1 ppm. The suspensions consisting of ca. 0.42 mol/L Fe³⁺ in 0.1 mol/L NaClO₄ were aged in quiescent conditions under a nitrogen atmosphere at room temperature. The suspension pH of the Ni ferrihydrite was 5.0 after 38 months aging, and that of the Fe ferrihydrite was 2.9 after 41 months.

Aliquots of the Fe ferrihydrite and Ni ferrihydrite were aged for different lengths of time: 8 days ("fresh") and 8, 17, 35, and 38 months for the Ni ferrihydrite coprecipitates, and 8 days ("fresh") and 17, 35, and 41 months for the Fe ferrihydrite. The "fresh" and aged ferrihydrite samples, with and without Ni, were transferred to a vacuum filtration flask, rinsed with deionized water, and air-dried. The effects, if any, of air-drying on mineral transformation were assumed to be minimal. The air-dried samples were used for powder X-ray diffraction (XRD), Mössbauer spectroscopy, and transmission electron microscopy (TEM).

Aliquots of the Ni ferrihydrite suspension were dissolved completely in concentrated Ultrex[®] HCl, and analyzed by inductively coupled plasma mass spectroscopy (ICP-MS, Hewlett Packard 4500). The [Fe]_{aq} and [Ni]_{aq} were determined simultaneously. These data allowed calculation of the Ni/(Ni + Fe) ratio or χ_{Ni} in the coprecipitates.

A subsample of the Ni ferrihydrite coprecipitate consisting of ~0.01 mol/L Fe³⁺ was treated with 100 mL of acidified 0.2-mol/L ammonium oxalate (pH 3.0 with HCl) solution for either 2 or 24 h at room temperature in light-excluded containers (Fey and LeRoux 1977). Following the desired extraction period, the suspension was centrifuged at 5000 rcf for 10 min. The supernatant was removed and filtered through a 0.2 µm filter, and analyzed by ICP-MS for Fe and Ni. The ammonium oxalate-extracted Ni ferrihydrite residue was washed twice with 0.1 mol/L NaClO₄ to remove dissolved metals, washed once with 0.1 mol/L NaHCO₃ to neutralize surface pH, and then air-dried. The air-dried sample was used for XRD and Mössbauer. An aliquot of the 41 month aged 2-line Fe ferrihydrite suspension was treated similarly. Details about the isolation and culturing of bacteria, media composition, biotransformation experiments, and analyses (XRD and Mössbauer) of the bioreduced samples have been described previously (Fredrickson et al. 2001).

METHODS

Powder XRD patterns and Mössbauer measurements were obtained as described previously (Fredrickson et al. 2001; Kukkadapu et al. 2001). The JADE+, V5 (Materials Data, Inc., Livermore, California) data analysis software package was used to subtract the XRD spectra of the ammonium-oxalate treated sample from the spectra of the untreated sample. All Mössbauer spectra were fitted (without thickness correction) with quadrupole splitting (QSDs) and/or hyperfine splitting distributions (HFDs) using the Voigt-based fitting method of Rancourt and Ping (1991). The following guidelines were used in the modeling of the Mössbauer data: (1) the Lorentzian half-width at half maximum (HWHM) was fixed at 0.097 mm/s, (2) all doublets were assumed to be symmetric, (3) for sextets, the ratio of the spectral areas of peak 1 to peak 3 was fixed at 3 and the ratio of the spectral areas of peak 2 to peak 3 was fixed at 2, and (4) no coupling of δ or center shift (CS) with the distributed hyperfine parameter (Δ or QS), or coupling of quadrupole shift parameter (ϵ) with the distributed hyperfine parameters (z), were allowed.

A drop of dilute, sonicated mineral suspension was applied directly to a 200-mesh copper grid coated with carbon-sputtered Formvar support film for TEM studies. Samples were examined on a JEOL 2010 HR TEM at 200 kV. Imaging and analyses were performed at electron-transparent edges of sample aggregates. Selected Area Electron Diffraction (SAED) was performed on the ferrihydrite. Ring patterns were obtained using the smallest SAED aperture at 40 s for d values smaller than 0.25 nm (errors ± 0.005 nm). Both images and SAED patterns were digitally collected, and were analyzed using Digital Micrograph and Desktop Microscopist software.

RESULTS

In the results that follow we describe our finding of 6-line ferrihydrite as an apparent transformation product of 2-line ferrihydrite in three different systems: (1) Ni ferrihydrite aged under aerobic conditions, (2) Fe ferrihydrite aged under aro-

bic conditions, and (3) Ni ferrihydrite incubated with an electron donor and an Fe-reducing bacterium under anoxic conditions. The intent of this presentation is to show that 6-line ferrihydrite was an unexpected, but common transformation product of 2-line ferrihydrite in suspensions that vary in pH, salt, minor element content, redox potential, and organic compounds.

Aerobic transformation of Ni ferrihydrite

Chemical analyses and powder XRD. The pH of the Ni ferrihydrite suspension decreased from pH 7.2 to pH 5.0 with 35 months of aging. Aging, however, had little effect on the mole fraction of Ni (χ_{Ni}) of the coprecipitate. The χ_{Ni} in the "fresh" sample was 0.038, whereas that of the 35 month aged sample was 0.035. The XRD patterns, on the other hand, changed significantly over the period of aging (Fig. 1). Two broad peaks, which are typical of 2-line ferrihydrite (Schwertmann and Cornell 1991), dominated the XRD patterns of the "fresh" (not shown), 8-, and 17 month aged samples (peaks at 0.22 to 0.25 and 0.15 nm). In addition to the 2-line ferrihydrite peaks, a sharper peak centered around 0.25 nm that is characteristic of the $d(110)$ peak of 6-line ferrihydrite (Eggleton and Fitzpatrick 1988) was apparent in the 8 month aged sample (Fig. 1a). The intensity of this peak ($d(110)$) increased after 17 months of aging (Fig. 1b). Other incipient peaks attributable to 6-line ferrihydrite were also evident in the 17 month spectrum (Eggleton and Fitzpatrick 1988; Carlson and Schwertmann 1981). Overall, the diffraction features in Figs. 1a and 1b implied the co-existence of 2-line ferrihydrite and 6-line ferrihydrite, with the latter increasing with aging time. The 6-line ferrihydrite contribution dominated the diffractogram of the 35 month sample (Fig. 1c). In addition to 6-line ferrihydrite, peaks assignable to goethite and hematite were also evident. The broad feature at $29^\circ 2\theta$ (0.31 nm) evident in the 8- and 17 month samples was noticeably absent in the 35 month sample, implying a decrease in the content of 2-line ferrihydrite.

Ammonium oxalate treatment. A 38 month aged Ni ferrihydrite, which displayed an XRD pattern identical to that of the 35 month sample (Fig. 2a; compare with Fig. 1c), was obtained after treatment with acid ammonium oxalate for 2 or 24 h. Both 2- and 6-line synthetic ferrihydrite dissolve quickly in acidified ammonium oxalate (Schwertmann 1959; Schwertmann and Fischer 1973; Eggleton and Fitzpatrick 1988). The treatment is used to separate poorly crystalline Fe^{3+} -oxides (ferrihydrites) from crystalline ones (goethite, hematite, or lepidocrocite) because the latter are less affected by the reagent (Schwertmann et al. 1982). Both the 2 and 24 h treatments dissolved similar amounts of Fe and Ni (75% of Fe and 65% of Ni for the 2 h treatment, and 80% Fe and 73% Ni for the 24 h treatment), implying that the mineral mixture consisted of 20 to 25% crystalline compounds.

The residual mineral fraction after oxalate extraction contained goethite and hematite (Fig. 2b). The difference pattern between the untreated and treated sample (Fig. 2c) displayed peaks that matched with those for 6-line ferrihydrite (Eggleton and Fitzpatrick 1988, Carlson and Schwertmann 1981) (Fig. 2c). Differential XRD following oxalate treatment is routinely performed to identify ferrihydrite in soils (Schulze 1994). The

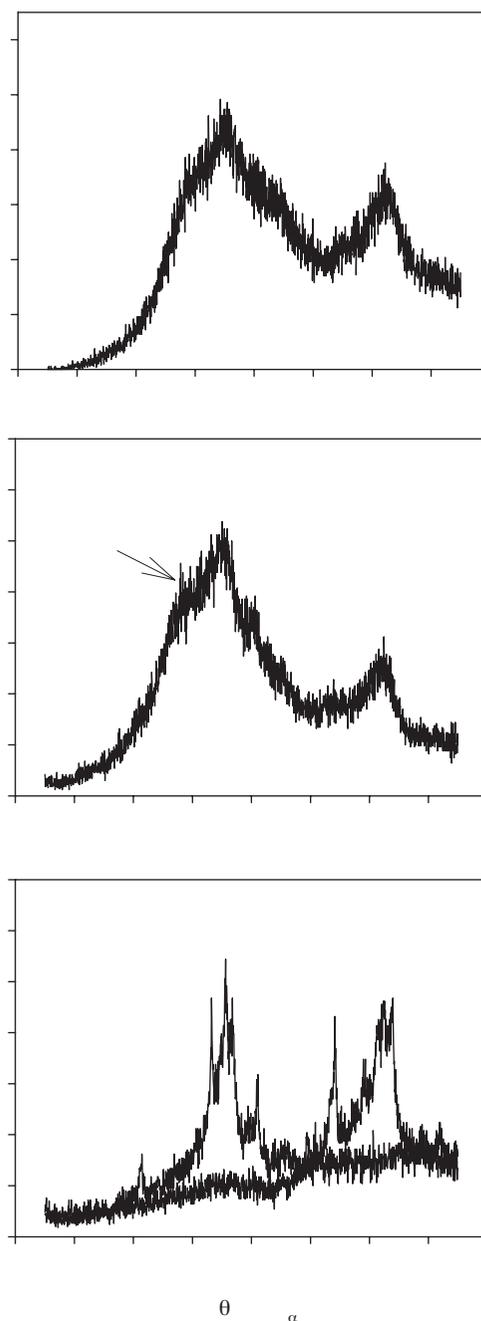


FIGURE 1. Powder XRD patterns of "aged" Ni ferrihydrite. (a) 8 months, (b) 17 months, and (c) 35 months. A diffractogram of a zero-background slide (ZBS) is included in Figure 1c. G and H in Figure 1c label goethite and hematite.

relative intensities of the peaks in Figure 2c, however, were different from those of typical 6-line ferrihydrite, possibly because the two broad peaks of residual 2-line ferrihydrite were buried under the 6-line pattern. The buried 2-line feature made it difficult to resolve the fifth and sixth lines of 6-line ferrihydrite. Eggleton and Fitzpatrick (1998) interpreted the asymmetry on the low-angle side of the $d(110)$ reflection (36.2°

$^{\circ}2\theta$) of 6-line ferrihydrite to result from 2-line ferrihydrite. The difference pattern in Figure 2c also displays the presence of two small, broad peaks at 0.45 and 0.32 nm. These peaks at 0.45 and 0.32 nm are not readily evident in Figure 2c, but are detectable from counts. Some believe (Towe and Bradley 1967; Eggleton and Fitzpatrick 1988) that they represent the $d(002)$ and $d(003)$ peaks of 6-line ferrihydrite. The first major peak of the 6-line pattern [$d(110)$] peak at $36^{\circ}2\theta$, or d -value of 0.25 nm] had a half-width at half maximum (HWHM) of $\sim 3^{\circ}2\theta$. The HWHM of the $d(110)$ line increases with decreasing crystallinity, and Carlson and Schwertmann (1981) have reported HWHM values in the range of 6 to $16^{\circ}2\theta$ for Si ferrihydrites.

Electron microscopy. The 35 month aged Ni ferrihydrite was analyzed by TEM to obtain microscopic evidence for the presence of both 2-line and 6-line ferrihydrite. In Figure 3a,

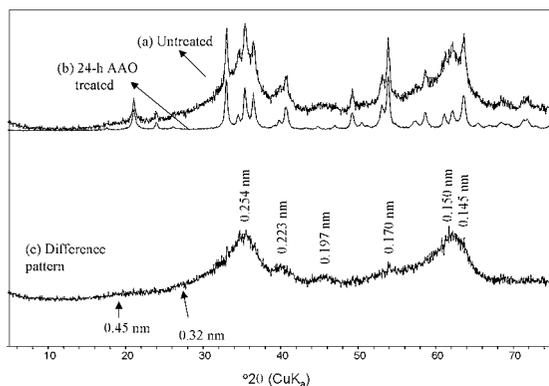
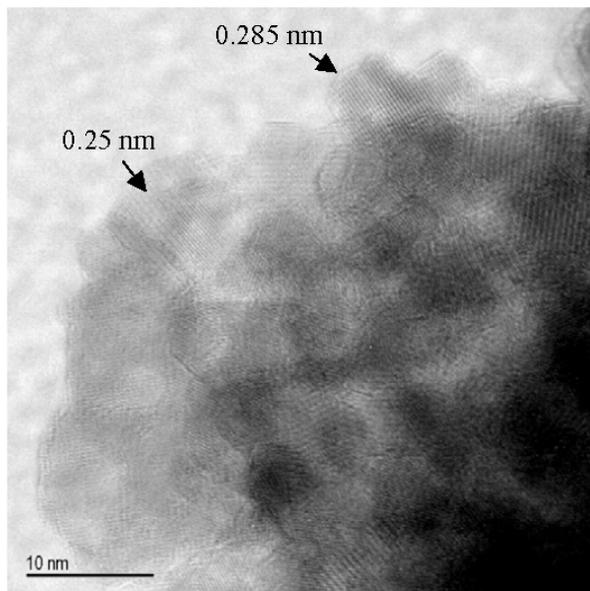


FIGURE 2. Powder XRD patterns of the 38 month aged Ni ferrihydrite before and after 24 h AAO treatment, and the difference pattern.

(a) Micrograph



pseudo-hexagonal crystallites with an average size of 6 to 10 nm were evident that were typical of 6-line ferrihydrite. The crystallites exhibited lattice fringes consistent with 6-line ferrihydrite at 0.25 nm and 0.285 nm. The crystallite morphology was similar to that of the 6-line ferrihydrite reported by Janney et al. (2000). An SAED pattern of these crystals (Fig. 3b) displayed d values (for the diffuse rings) that were within errors of determination from those reported for 6-line ferrihydrite (Janney et al. 2000).

Selected areas in the 35 month Ni ferrihydrite also showed the presence of 2-line ferrihydrite (Figs. 4a and 4b). The crystallites were 2 to 4 times smaller than those of 6-line Ni ferrihydrite (Fig. 4a) and did not exhibit coherent domains. The SAED pattern displayed two bright rings with shoulders on each side, identical to the 2-line ferrihydrite of Janney et al. (2000). In addition to the 2- and 6-line ferrihydrite, the micrographs displayed regions that were rich in acicular goethite crystals (~ 30 nm long, top right corner, Fig. 4a) and hexagonally shaped hematite (not shown).

Mössbauer spectroscopy. Room-temperature Mössbauer measurements, alone, are not adequate to characterize crystallinity of ferrihydrite (Murad 1996) because 2-line ferrihydrite exhibits a doublet feature both at room temperature and 77 K, whereas 6-line ferrihydrite exhibits a doublet only at room temperature. A “collapsed” sextet dominates the spectrum of 6-line ferrihydrite at 77 K (Murad 1996). The presence of a doublet in the room temperature and 77 K spectra of the “fresh” Ni ferrihydrite (Figs. 5a and 5d) indicated that it was the 2-line form.

A “collapsed” sextet feature that is characteristic of 6-line ferrihydrite, however, was evident in the 77 K spectra of the 17- and 35 month aged samples (Figs. 5e and 5f, black trace),

(b) SAED pattern

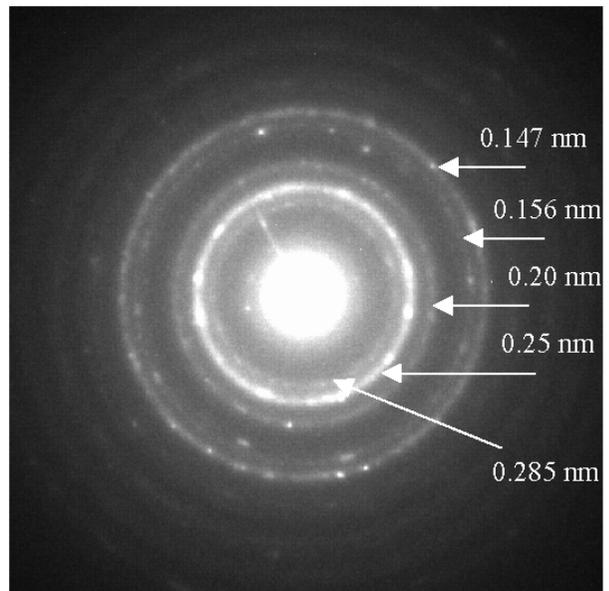


FIGURE 3. TEM of 35 month aged Ni ferrihydrite: (a) hexagonal crystals of 6-line Ni ferrihydrite that range from 6 to 10 nm, and lattice fringes and (b) SAED with d -values of rings.

in agreement with XRD observations (Fig. 1). The 8 month sample exhibited spectra similar to those of the 17 month sample (not shown). Note that in the 17 month sample, only a doublet feature was evident in the room temperature spectrum (Fig. 5b). In addition to the “collapsed” sextet feature (6-line ferrihydrite), two minor sextets attributable to goethite (inner sextet) and hematite (outer sextet) were also present in the 77 K spectra of the aged samples. The contribution of both goethite and hematite to the spectra was greater in the 35 month sample (Fig. 5f, black trace). The presence of hematite and goethite in the 35 month was also evident in room temperature spectrum of the 35 month aged sample (* and + in Fig. 5c). A comparison of the Mössbauer spectrum of the 24 h AAO treated 35 month sample with the untreated one (Fig. 5f) revealed that (1) the bulk of the 35 month aged sample (untreated) was 6-line ferrihydrite (in accordance with XRD), (2) AAO preferentially dissolved 6-line ferrihydrite without significantly affecting the crystallinity of goethite and hematite, and (3) the Mössbauer signature of goethite and hematite in the 24 h AAO sample was very similar to the crystalline Fe^{3+} oxide fraction in the untreated sample.

A two-Gaussian QSD fit was used to model the room temperature and 77 K spectra of the “fresh” sample (Figs. 5a and 5d). The QS of maximum probability deduced from the QSD corresponds to the average Fe^{3+} distortion and the width of the distribution corresponds to the variation in Fe^{3+} site geometry (Rancourt 1988; Murad 1996; Rancourt et al. 1994). The derived Mössbauer parameters are shown in Table 1. The room temperature parameters were similar to those recently reported for a room temperature spectrum of a synthetic 2-line ferrihydrite, and are typical of Fe^{3+} in octahedral coordination (Rancourt et al. 2001). The assignment of Fe^{3+} sites to octahe-

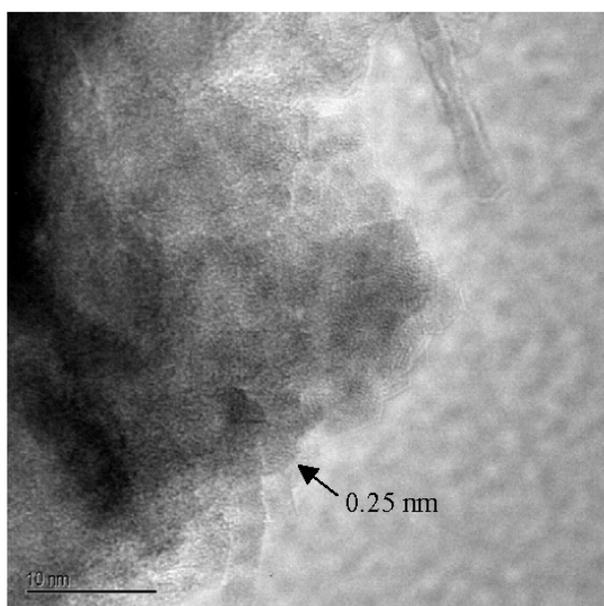
dral coordination is in agreement with the Mössbauer spectroscopy data of ferrihydrite by Pankhurst and Pollard (1992). The objective of the modeling of the 77 K spectrum was to utilize its fit parameters in describing the 35 month aged sample (discussed below).

A combined QSD + HFD model was used to fit the 77 K spectra of the 35 month aged samples (untreated and 24 h AAO; Fig. 6, Table 1). In these, it was possible to resolve the contributions from various sites with certain assumptions. The objective of the spectral modeling of the AAO-treated sample (Fig. 6a) was to identify the Mössbauer parameters of Ni goethite and Ni hematite that could be used along with those for 2-line Ni ferrihydrite to qualitatively estimate (assuming identical f -factors for all phases) the contribution of 6-line ferrihydrite to the untreated sample (Fig. 6b). A two-Gaussian component distribution was used to model the 6-line ferrihydrite contribution (Fig. 6c; Table 1). The derived relative areas of goethite and hematite together were 20% in the 35 month aged Ni ferrihydrite, a value that agreed well with the crystalline Fe(III) oxide content estimated from AAO treatment. The computed spectral areas of 2-line and 6-line ferrihydrite were 15 and 65%, respectively, in the same sample. The $\langle e \rangle$ value of 0.21 in for hematite indicated that the phase was highly crystalline (Dang et al. 1998).

Aerobic transformation of Fe ferrihydrite

The pH of the Fe ferrihydrite suspension decreased more with aging than did the Ni ferrihydrite suspension. The pH of the “fresh” Fe ferrihydrite suspension was 7.2, and that of the 41 month aged suspension was 2.9. Powder XRD of the “fresh” sample revealed the presence of two broad peaks (0.22 to 0.25 and 0.15 nm) characteristic of 2-line ferrihydrite (Fig. 7a). The

(a) Micrograph



(b) SAED pattern

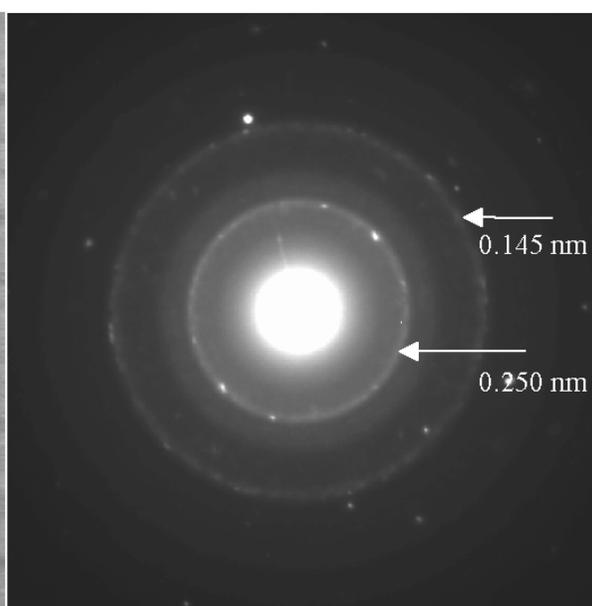


FIGURE 4. TEM of 35 month aged Ni ferrihydrite: (a) small crystals of 2-line Ni ferrihydrite (up to 4 nm) and (b) SAED with d -values of rings.

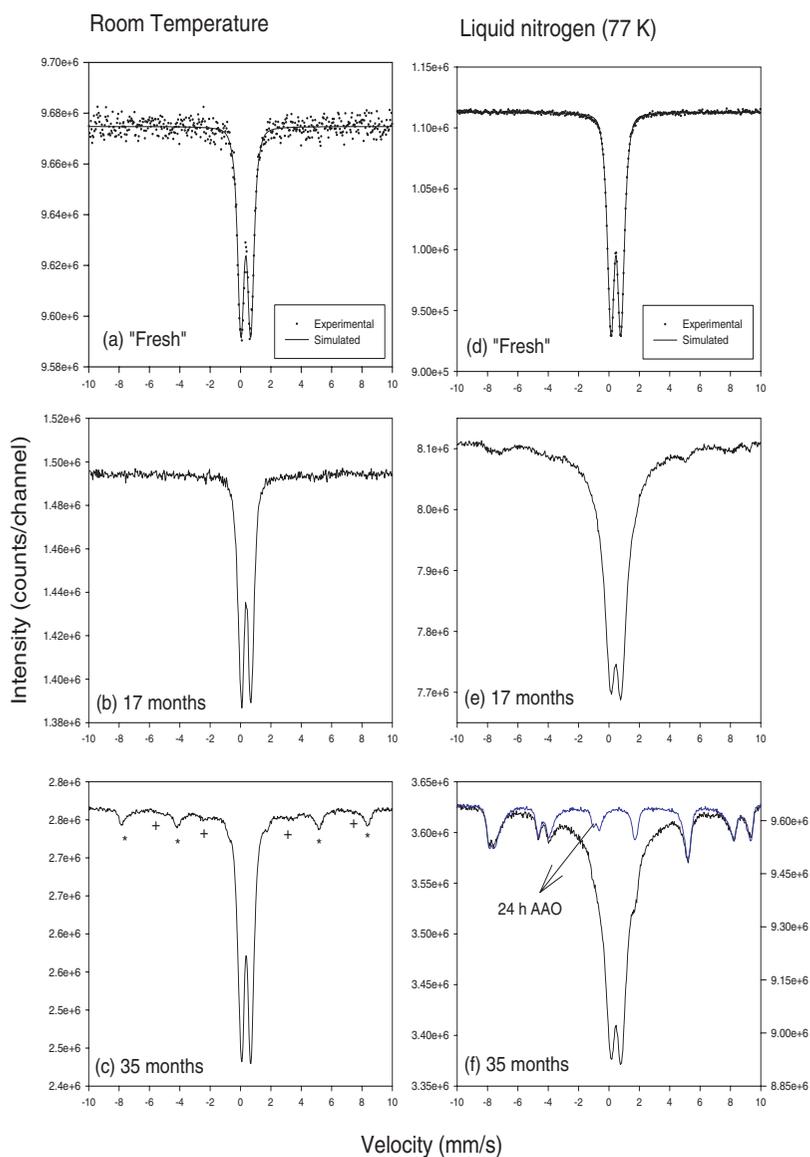


FIGURE 5. Room-temperature (a–c) and liquid nitrogen (d–f) Mössbauer spectra of “fresh” and aged Ni ferrihydrite samples. The Mössbauer spectrum of the 24 h AAO extracted, 35-month sample is shown in Figure 5f (in gray).

TABLE 1. Mössbauer parameters for various ferrihydrites and reacted ferrihydrite samples

Sample	Fe-mineral	$\langle CS \rangle$ (mm/s)	$\langle QS \rangle$ (mm/s)	$\langle \epsilon \rangle$ (mm/s)	$\langle B_{hf} \rangle$ (Tesla)
Aerobic - Ni ferrihydrite					
“Fresh” (RT- spectrum)	2-line ferrihydrite	0.34	0.71	–	–
“Fresh” (77 K spectrum)	2-line ferrihydrite	0.46	0.72	–	–
24-h AAO/35-month (77 K)	Ni goethite	0.49	–	–0.1	47.6
	Ni hematite	0.52	–	0.21	53
35-month (untreated) (77 K)	Ni goethite	0.49	–	–0.1	47.6
	Ni hematite	0.52	–	0.21	53
	2-line ferrihydrite	0.46	0.72	–	–
	6-line ferrihydrite	0.46	–	–0.006	13.4
Aerobic - Fe ferrihydrite					
41-month (untreated) (77 K)	goethite	0.47	–	–0.13	49
	6-line ferrihydrite	0.43	–	–0.02	21.5
Anaerobic - Ni ferrihydrite					
Bioreduced in HCO_3^- buffer (77 K)	goethite	0.49	–	–0.11	47.3
	2-line ferrihydrite	0.45	0.73	–	–
	6-line ferrihydrite	0.46	–	–0.006	16.61

Note: $\langle CS \rangle$ = average center shift; $\langle QS \rangle$ = average quadrupole shift; $\langle \epsilon \rangle$ = average quadrupole shift parameter and $\langle B_{hf} \rangle$ = average hyperfine field.

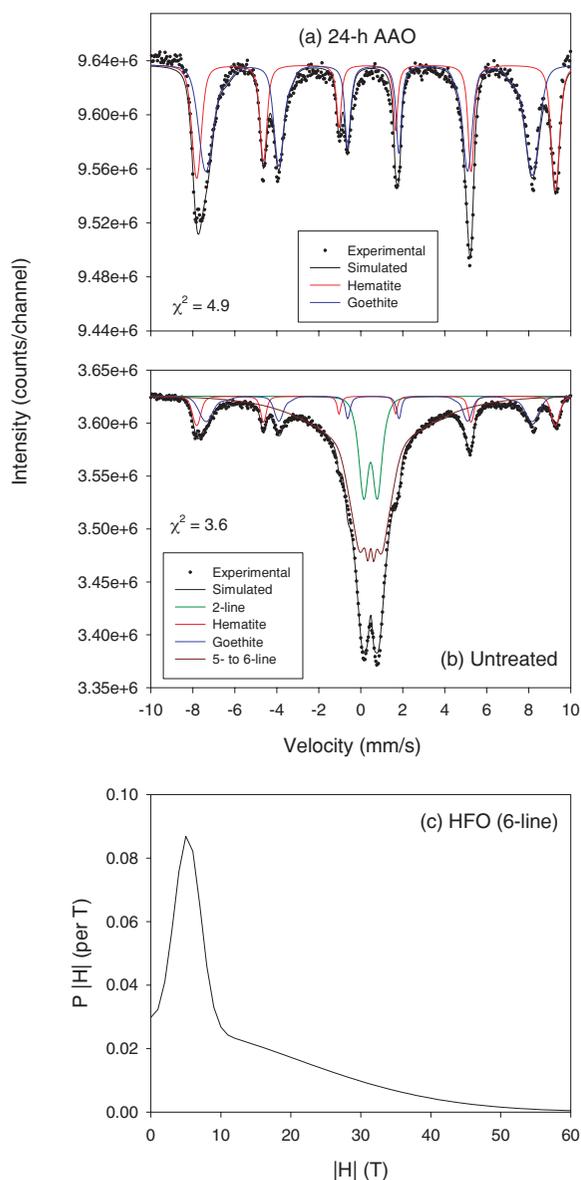


FIGURE 6. Experimental and simulated 77 K Mössbauer spectra of Ni ferrihydrite (a) the 24 h AAO/35-month aged sample, (b) the untreated 35-month sample, (c) HFD distribution of 6-line component of 77 K 35 month sample.

additional broad peak at 0.25 nm in the spectrum was due to glycerol; this particular diffraction pattern was the only one collected in presence of glycerol. The 17 month sample differed considerably in its diffraction pattern from Ni ferrihydrite (Fig. 7b, compare with Fig. 1b). Sharp peaks that were characteristic of goethite dominated its spectrum (G in Fig. 7b), in contrast to the Ni ferrihydrite. Diffraction maxima due to 6-line ferrihydrite that dominated in the Ni ferrihydrite were far less prominent. The goethite content increased significantly with continued aging. Goethite dominated the spectrum of the 35 month (Fig. 7c) and 41 month (Fig. 8a) samples.

Acid ammonium oxalate extraction was carried out on the

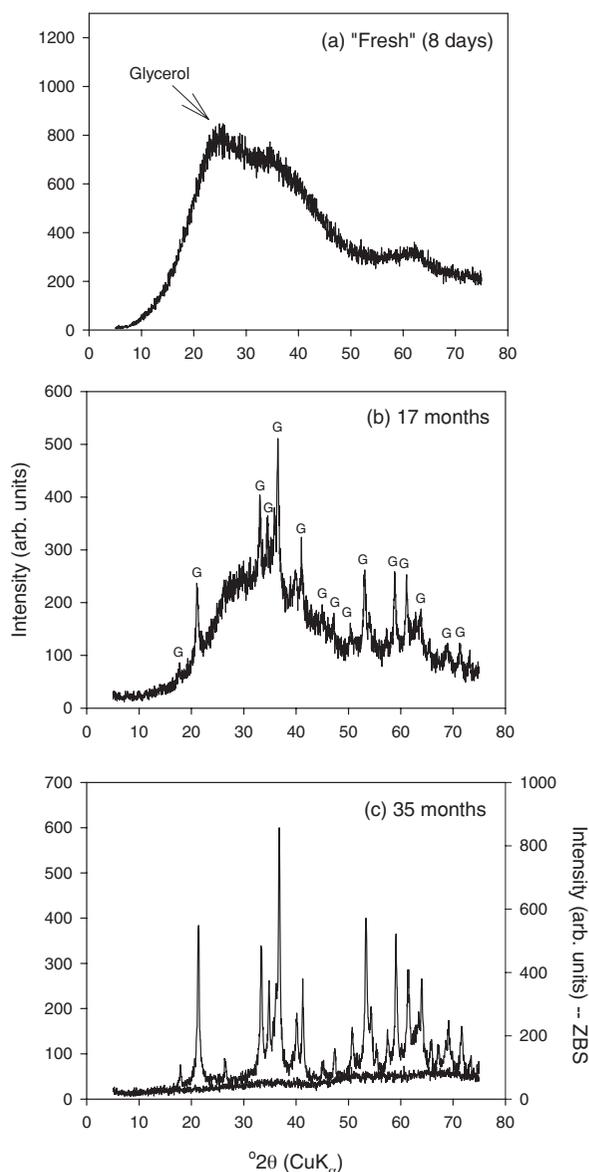


FIGURE 7. Powder XRD patterns of "aged" Fe ferrihydrite: (a) "fresh" (8 days), (b) 17 months, and (c) 35 months. G in Figure 7b labels goethite. The peak at 26 °2θ in the "fresh" sample is from glycerol.

41 month sample (Fig. 8a) to identify the nature of the residual ferrihydrite. Approximately 33 and 41% of Fe was dissolved in 2 and 24 h treatments, a value that was considerably less than for the 38 month aged Ni ferrihydrite (~75%). An XRD spectrum of the 24 h ammonium oxalate-treated sample displayed peaks primarily from goethite (Fig. 8b). A comparison of the untreated and oxalate-treated samples revealed that the crystallinity (based on line widths) of the goethite was unaffected by AAO treatment. The difference pattern between the untreated and treated samples displayed peaks (Fig. 8c) that matched well with those of 6-line ferrihydrite (Eggleton and Fitzpatrick 1988). Little, if any contribution of 2-line

ferrihydrate was noted to the difference pattern.

Mössbauer spectroscopy. Room temperature and 77 K Mössbauer measurements of the 41 month aged ferrihydrate displayed the presence of residual ferrihydrate that was removed by AAO treatment (Fig. 9). The contribution of this phase to the 77 K spectrum (Fig. 9c) was described with a collapsed sextet, consistent with 6-line ferrihydrate, and the noted HFD distribution (Fig. 9d). Two-line ferrihydrate was not included in the fit because its presence was not indicated by XRD. The sextet was assigned to goethite, based on XRD data. As noted for Ni ferrihydrate, AAO removed 6-line ferrihydrate (the collapsed sextet feature in the 77 K spectrum) without significantly affecting the crystallinity of goethite. The Mössbauer signature of goethite in the AAO-treated sample was virtually identical to that of the crystalline Fe^{3+} oxide fraction in the untreated sample. The differences in the HFD distributions for 6-line ferrihydrate in the Fe ferrihydrate and Ni ferrihydrate samples were attributed to variation in the superparamagnetic behavior of these solids. Coprecipitated metal ions affect superparamagnetic character of ferrihydrate by decreasing interactions between surface unpaired Fe^{3+} spins of the ferrihydrate particles (Zhao et al. 1996).

Bioreductive transformation of Ni ferrihydrate

Recently, we studied the biotransformation of 14 month aged Ni ferrihydrate by *Shewanella putrefaciens*, strain CN32, under anoxic conditions (Fredrickson et al. 2001). The starting material was primarily 2-line ferrihydrate, but showed incipient 6-line diffraction features like those of the 17 month aged Ni ferrihydrate (Fig. 1b). Incubations with the 14 month Ni ferrihydrate (0.05 mol/L Fe^{3+}) and CN32 were carried out for 32 days at circumneutral pH with lactate (0.03 mol/L) as the electron donor under various media and buffer conditions (0.03 mol/L bicarbonate or PIPES). CN32 oxidizes lactate to acetate and bicarbonate, and disposes of electron equivalents to ferrihydrate, the electron acceptor. This microbiologically mediated electron transfer reaction causes reduction of Fe^{3+} to Fe^{2+} .

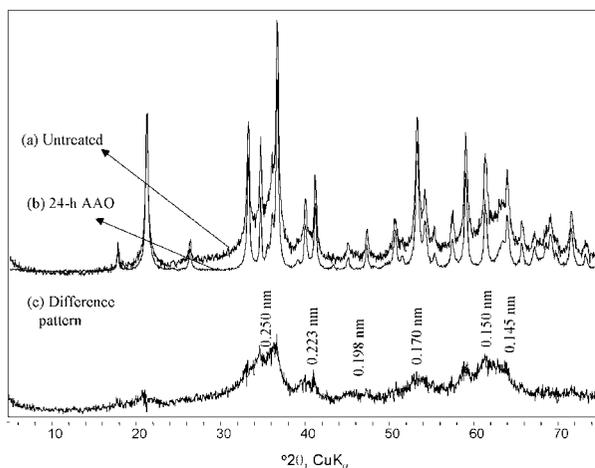


FIGURE 8. Powder XRD patterns of 41 month aged Fe ferrihydrate before and after 24 hour AAO treatment, and the difference pattern. Peaks other than 6-line in the difference pattern are artifacts of the subtraction procedure.

The Fe^{2+} so generated is reactive with the residual ferrihydrate (Zachara et al. 2002). A variety of mineral conversion products was observed, and the products were most dependent on the presence or absence of phosphate (P) or anthraquinone disulfonate (AQDS), a quinone that facilitates ferrihydrate reduction (Nevin and Lovley 2000). These mineral transformation products included magnetite, siderite, vivianite, goethite, and 6-line ferrihydrate.

In Fredrickson et al. (2001), limited discussion was given to the formation of 6-line ferrihydrate because the paper had other objectives. One XRD analysis was presented from the incubation that yielded 6-line ferrihydrate as a primary biotransformation product (bicarbonate, no AQDS, no P). Here we utilize newly collected 77 K Mössbauer data (bicarbonate, no AQDS, no P) from that phase association (Fig. 10) to draw comparisons to those obtained through aerobic ferrihydrate aging presented in previous sections. Additionally, Mössbauer data are presented for a mineral residue from another incubation in that series (PIPES, AQDS, no P) that yielded magnetite as a primary biotransformation product in association with residual ferrihydrate (Fig. 11). Our intent in presenting these unpublished results is to document the apparent consistent transformation of 2-line to 6-line ferrihydrate under both aerobic and anoxic conditions.

The 77 K Mössbauer spectrum of the mineral residue from the bicarbonate, no AQDS, no P incubation (Fig. 10) showed distinct similarities to those observed for the 35 month Ni ferrihydrate (Fig. 5) and the 41 month Fe ferrihydrate (Fig. 9). All contained sextets associated with crystalline Fe^{3+} oxide (goethite primarily, with minor hematite), and a broadened central doublet, or collapsed sextet (superparamagnetic component) believed to represent 6-line ferrihydrate. What differed between samples were the apparent mass distributions of the individual phases as implied by spectral areas. The difference in the 6-line ferrihydrate HFD distribution (Fig. 10b) from that of the aerobic samples was attributed to superparamagnetic behavioral effects, due to different media composition, redox potential, and organic contents of the aerobic and anaerobic incubations. This particular incubation with CN32 produced the lowest levels of Fe^{3+} reduction of all those studied by Fredrickson et al. (2001), and contained <5% Fe^{2+} . This sorbed Fe^{2+} was barely visible by Mössbauer spectroscopy (* in Fig. 10, the high field component). The microbiologically incubated Ni ferrihydrate displayed significantly more transformation of 2-line ferrihydrate to 6-line ferrihydrate and goethite than did the aerobically aged Ni ferrihydrate, in spite of higher pH and a much shorter overall equilibration time (e.g., 32 days vs. 15 months).

A common observation in studies of the microbiological transformation of ferrihydrate is the presence of residual ferrihydrate along with more crystalline, biogenic phases (Fredrickson et al. 1998; Fredrickson et al. 2001; Zachara et al. 2002). The crystalline character of this residual ferrihydrate has not been determined. Room-temperature Mössbauer measurements on the mineral residue from the PIPES, AQDS, no P incubation with Ni ferrihydrate (Fredrickson et al. 2001) showed the presence of magnetite (sextets) and residual ferrihydrate (central doublet) in the 32 day mineral residue (Fig. 11a). The biogenic material, that exhibited a $\text{Fe}^{2+}/\text{Fe}^{3+}$ ratio of 0.2, dis-

played a Mössbauer spectrum similar to that of abiotically synthesized magnetite with $\text{Fe}^{2+}/\text{Fe}^{3+} = 0.25$ (Jolivet et al. 1992). The central doublet disappeared upon cooling the bioreduced sample to 77 K (Fig. 11b), indicating that all of the residual ferrihydrite associated with the magnetite was more crystalline than the 2-line ferrihydrite (note Mössbauer response of 2-line ferrihydrite to 77 K, Fig. 5d). The crystalline ferrihydrite in the magnetite-containing sample exhibited a different response to 77 K than did that in all the other samples, where line broadening and partial collapse were the primary responses to cooling (e.g., Figs. 5e and 5f, 9b, 10a). The crystalline ferrihydrite that was present with magnetite apparently exhibited a different magnetic ordering temperature than did 6-line ferrihydrite that formed at lower Fe^{2+} concentrations.

DISCUSSION

Formation of 6-line ferrihydrite under oxic conditions

2-line Ni ferrihydrite. The transformation of freshly precipitated 2-line Ni ferrihydrite to goethite and/or hematite was slow under the conditions of this study. Past research (Cornell and Schwertmann 1996) has shown that coprecipitated metal cations (e.g., Co, Ni) retard ferrihydrite crystallization. Most studies of ferrihydrite crystallization have been carried out in more dilute ferrihydrite suspensions at higher temperature (Cornell et al. 1992; Ford et al. 1997; Ford et al. 1999) and higher pH (Cornell et al. 1992). Transformation rates are more rapid under such conditions. For example, at 70 °C and pH 12, ~40 % of Ni ferrihydrite crystallized in one day (Cornell et al. 1992), making it difficult to identify changes in ferrihydrite structure. Following such transformations at lower temperature, as was done here, provided an opportunity to follow progressive changes in structure (ordering) of the Ni ferrihydrite coprecipitate. In addition, our use of concentrated, initially circumneutral ferrihydrite suspensions (close to the pH of zero net charge of ferrihydrite), encouraged the initial aggregation of ferrihydrite particles (Schwertmann et al. 1999, Banfield et al. 2000) that appears necessary for thermodynamic transformation to more stable hematite (Schwertmann et al. 1999) and possibly goethite (Banfield et al. 2000).

Structural changes in the 2-line Ni ferrihydrite were clearly evident as aging proceeded (Figs. 1 and 5). Partial conversion of 2-line Ni ferrihydrite to 6-line Ni ferrihydrite occurred over the first 17 months of aging (Figs. 1 and 5). The presence of a “collapsed” sextet in the 17 month 77 K Mössbauer spectrum, not present in the “fresh” sample, demonstrated such a change. We attribute the “collapsed” sextet to 6-line ferrihydrite because goethite and hematite were not evident in the XRD pattern (in which the XRD detection limit is approximately 5 mass%). With further aging, the 6-line Ni ferrihydrite content increased considerably. The increase in 6-line Ni ferrihydrite content was clearly evident from comparison of the XRD spectra of the 17 month (Fig. 1b) and 38 month (Fig. 2c) samples. Mössbauer modeling (Fig. 6b) showed that 6-line Ni ferrihydrite accounted for approximately 65 mass percent of the 35 month sample.

The increase in 6-line Ni ferrihydrite content was paralleled by small increases in goethite and hematite. Goethite and hematite increased from a combined percentage of approximately

5% in the 17 month sample to 22% in the 35 month sample. The goethite-to-hematite ratio in the 35 month sample was 1:1 whereas the ratio was 4:1 for a dilute 2-line Fe ferrihydrite suspension (0.025 mol/L) aged for 15 months at a similar pH and temperature (Schwertmann and Murad 1983). The transformation of 2-line to 6-line ferrihydrite, formation of only small amounts of goethite and hematite, and low goethite-to-hematite ratio, collectively indicate that aggregation-based crystal growth (crystallization within ferrihydrite aggregates) was responsible for 6-line ferrihydrite formation. The high suspension density of ferrihydrite and the moderate ionic strength used in our experiments may have been important in enhancing aggregation. Our study could not discriminate whether hematite was formed from 2-line ferrihydrite directly or via 6-line ferrihydrite. Both 2-line ferrihydrite and 6-line ferrihydrite can transform to hematite under various conditions (Schwertmann et al. 1999).

2-line Fe ferrihydrite. Aging experiments were performed with 2-line Fe ferrihydrite to identify the impacts of Ni, by comparison, on ferrihydrite transformation. Schwertmann and Murad (1983) aged 2-line Fe ferrihydrite at 24 °C for 31 months and characterized the transformation products. Our study differed from the Schwertmann and Murad (1983) experiments in that: (1) a higher suspension density of ferrihydrite was used here (0.42 mol/L vs. 0.025 mol/L), (2) the suspension pH was not adjusted, and (3) the ferrihydrite was aged in 0.1 mol/L NaClO_4 . The pH of our “fresh” suspension was 7.2 and that of the 41 month aged sample was 2.9. In contrast, Schwertmann and Murad (1983) maintained a nearly constant suspension pH. The low final pH of our suspensions strongly encouraged ferrihydrite dissolution over the later stages of aging.

The aging of freshly precipitated 2-line Fe ferrihydrite yielded a different product suite than did that of the 2-line Ni ferrihydrite. Goethite was generated in larger quantities than was 6-line ferrihydrite (Figs. 7b and 7c) in the Fe ferrihydrite suspension, and hematite was present in small concentrations (Fig. 9c). The formation of goethite increased with aging duration and lower suspension pH. The enhanced formation of goethite indicated that dissolution of 2-line ferrihydrite and solution phase precipitation (Feitknecht and Michaelis 1962) was dominant over aggregation-induced crystallization. Apparently, the presence of coprecipitated Ni^{2+} in the Ni ferrihydrite moderated pH change and retarded the dissolution process.

Our finding of appreciable amounts of 6-line ferrihydrite was unexpected because no such phase was observed in the study of Schwertmann and Murad (1983), wherein the conversion products were goethite and hematite in various ratios. The goethite-to-hematite ratio increased from 0.33 to 7.3 as the pH of the suspension decreased from 7 to 3, the pH range of the present study. Apparently, the high degree of ferrihydrite conversion to goethite and hematite observed by Schwertmann and Murad (81 to 96%) perhaps precluded the detection, if any, of 6-line ferrihydrite.

Aggregation/crystallization effects. We observed significant differences in suspension pH and mineralogical products after aging 2-line Fe and Ni ferrihydrites under aerobic conditions. The stoichiometric conversion of ferrihydrite (approximated as $\text{Fe}(\text{OH})_3$ for simplicity) to goethite or hematite

involves no pH change:



The proton condition of the 2-line ferrihydrite suspension at the initiation of the experiment (pH 7) may be defined as:

$$\text{TOT}_\text{H} = 2[\text{FeOH}_2^+]\text{s} + [\text{FeOH}]\text{s} + [\text{H}^+]\text{aq} - [\text{FeO}^-]\text{s} - [\text{OH}^-]\text{aq} \quad (3)$$

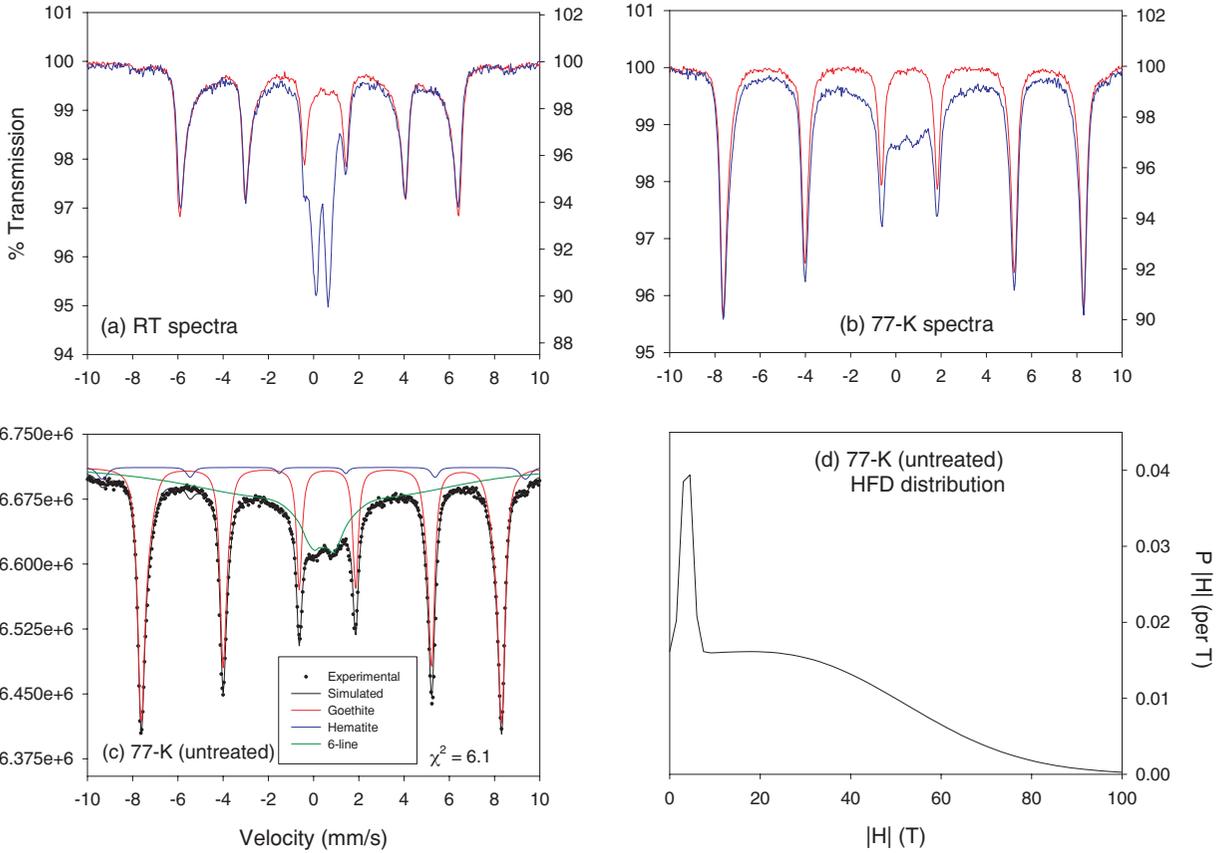


FIGURE 9. Room temperature (a) and liquid nitrogen (b) Mössbauer spectra of 41 month aged samples, the untreated sample is recorded in blue and the 24 h AAO treated sample in red. (c) Spectral modeling of the 77 K Mössbauer spectrum of the untreated sample, and (d) the HFD distribution used for 6-line ferrihydrite.

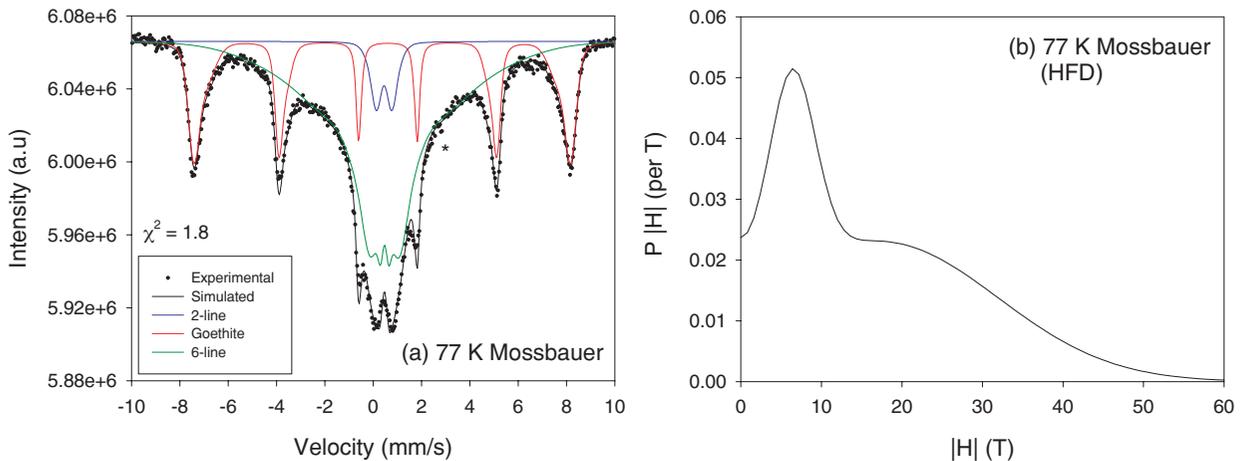
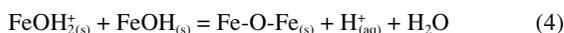


FIGURE 10. Mineral products resulting from the incubation of Ni ferrihydrite with CN32 in bicarbonate buffer without P and AQDS. (a) Experimental and simulated 77 K Mössbauer spectra and (b) HFD distribution of 6-line ferrihydrite.

where $[\text{FeOH}_2^+]_s$, $[\text{FeOH}]_s$, and $[\text{FeO}]_s$ are hydroxylated surface sites on ferrihydrite. The aging and recrystallization of 2-line ferrihydrite involves a decrease in surface area, and an increase in structural Fe-O-Fe bonds through condensation of hydroxylated surface sites that are in relatively high concentration (0.2 mol of FeOH/mole of Fe, Dzombak and Morel 1990). Because the initial pH of the aging solution (pH = 7.2) was below that of the PZC (pH ~8) of the oxide surface (e.g., $[\text{FeOH}_2^+]_s > [\text{FeO}]_s$), surface-site condensation leads to acidification of the recrystallizing mineral suspension:



Proton release in our experiments was particularly pronounced because of (1) the high suspension concentration of 2-line ferrihydrite (0.42 mol/L), and concomitant high concentration of FeOH_2^+ , and (2) the low buffering capacity of the suspension provided only by surface reactions of the 2-line ferrihydrite and oxide conversion products of lower surface area and site density. This pH decrease that was promoted by both 6-line ferrihydrite and goethite formation further encouraged the conversion to goethite by enhancing the solubility and dissolution of ferrihydrite. A sizable pH decrease was also observed in the aging/transformation of 2-line ferrihydrite to hematite at 92 °C by Combes et al. (1990), who termed the process

deprotonation-oxalation.

Coprecipitated Ni retarded pH change and conversion to the more crystalline phases goethite and hematite. We speculate that coprecipitated Ni occupied hydroxylated surface sites on the ferrihydrite:



The proton evolved in (5) was neutralized during the hydrolytic synthesis of HFO. This coprecipitated Ni was present in a concentration that would saturate approximately 25% of the surface sites. Coprecipitated Ni retarded the progress of reaction (4) by decreasing the surface concentrations of $\text{FeOH}_{2(s)}^+$ and $\text{FeOH}_{(s)}$ through a combination of mass action and electrostatic effects. The net result was to reduce the formation of new Fe-O-Fe bonds, to decrease the extent of pH change, and to minimize goethite formation under aerobic conditions.

Formation of 6-line/“well” organized ferrihydrite under anaerobic conditions

The observation of rapid formation of 6-line ferrihydrite and goethite from Ni ferrihydrite under anaerobic, ferruginic conditions as observed in the bicarbonate, no AQDS, no P incubation (Fig. 10) was a curious one. In no other biotransformation study performed by our group [e.g., Fredrickson et al. 1998; Fredrickson et al. 2001; Zachara et al. 2002; Fredrickson et al. (2003)] has 6-line ferrihydrite appeared in such high concentration as a product phase. Previously, we have observed rapid recrystallization of 2-line ferrihydrite to hematite and goethite under Fe^{3+} -reducing conditions promoted by CN32 at circumneutral pH when the electron donor concentration was low (0.5 mmol/L) and Fe^{3+} suspension density was high (150 mmol/L) (Zachara et al. 2002). These conditions promoted the slow and incomplete reduction of ferrihydrite. Goethite formation was attributed to the sorption of biogenic Fe^{2+} on aggregated ferrihydrite, which appeared to greatly catalyze recrystallization. Such catalytic effects have been observed by others in abiotic, Fe^{2+} -spiked systems (Fischer 1972; Andreeva et al. 1995; Fredrickson et al. 2003), but the mechanisms involved have not been adequately explained. The addition of CN32 (ca $>10^6$ cells/mL) to ferrihydrite suspensions at pH 7 induces rapid and complete flocculation of the ferrihydrite, possibly implying a greater degree of aggregation of the ferrihydrite nano-particles in microbe-containing suspensions. The significant presence of 6-line ferrihydrite in Figure 10 is partially explained by the effect of Ni, which retards both recrystallization (Cornell et al. 1992) and bioreduction (Fredrickson et al. 2001), and possibly favors 6-line ferrihydrite formation through indirect or competitive kinetic effects.

The absence of the doublet in the 77 K spectrum in the magnetite-containing sample (Fig. 11b) implied that the residual ferrihydrite was significantly different from the 2-line form (Fig. 5d). This residual phase is more crystalline/organized than 2-line ferrihydrite. This “well” organized ferrihydrite may be an intermediate phase in the transformation of ferrihydrite to magnetite, or it may be an alternative product resulting from the reaction of Fe^{2+} with ferrihydrite (e.g., Fischer 1972; Fredrickson et al. 2003; Ardizzone and Formaro 1983). The

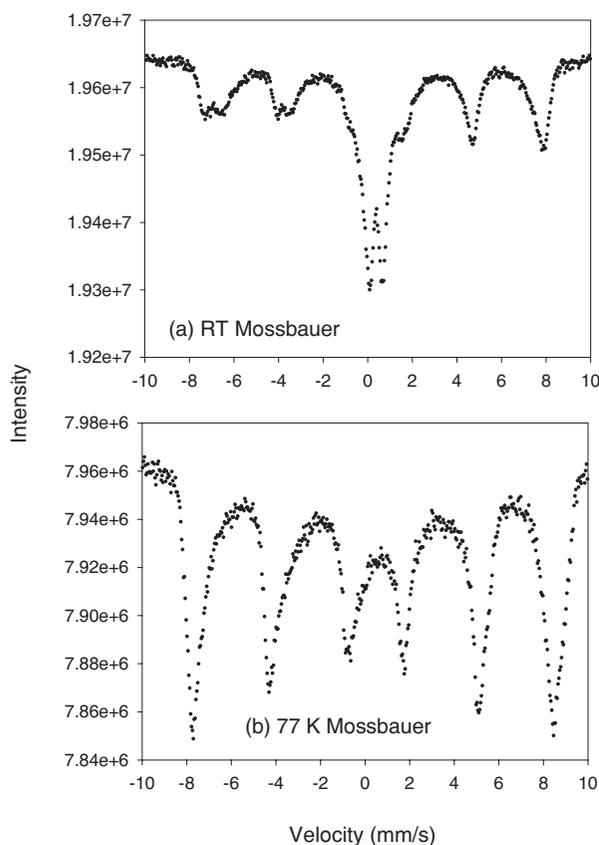


FIGURE 11. Mineral products resulting from the incubation of Ni ferrihydrite with CN32 in PIPES buffer with AQDS (a) Room temperature Mössbauer spectrum and (b) 77 K Mössbauer spectrum.

rapid formation (in days) of 6-line ferrihydrite and magnetite (in the presence of AQDS and CN32; Fredrickson et al. 2001), as compared to months in abiotic studies suggests that the biogenic Fe^{2+} flux rate accelerated 2-line ferrihydrite transformation to 6-line ferrihydrite.

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